**Aarambh Classes**

**Class IX**

**Chemistry Worksheet**

**IS MATTER AROUND US PURE**

1. Find the mass of potassium chloride needed to form a saturated solution in 60 g of water at 303K ? Given that solubility of the KCl is 37/100 g at this temperature .
2. What is the mass of sodium chloride neededto form a saturated solution in 50 g of water at 30C ?Solubility of sodium chloride is 36 g at 30C ?
3. A solution was prepared by dissolving 25 g of sugar in 100 g of water .Calculate the concentration of the solution .
4. 16 g of NaOH is dissolved in 100 g of water at 25C to form a saturated solution .Find the mass percentage or concentration of solute in solution .
5. A solution is made from 35 ml of methanol and 65 ml of water .Calculate the concentration of the solution .
6. Calculate the volume of ethanol in 200 ml aqueous solution if the concentration of solution is 20% .
7. The solubility of ammonium chloride in water at various temperatures is given below :

Temperature : 10C 20C 40C 60C 80C

Solubility : 24 g 37 g 41 g 55 g 66 g

What mass of ammonium chloride would be needed to make a sataurated solution of ammonium chloride in fifty grams of water at 40C ?

1. 21.5 g of sodium chloride dissolves in 60 g of water at 25C .Calculate the solubility of sodium chloride in water ta that temperature .
2. 9.72 g of potassium chloride dissolves in 30 g of water at 70C .Calculate the solubility of potassium chloride at that temperature .
3. A solution contains 5.6 ml of alcohol mixed with 75 ml of water .Calculate the concentration of this solution .